

## UNIT : 19 PURIFICATION AND CHARACTERISATION OF ORGANIC COMPOUNDS

### Important Points

1. **Common techniques of purification of organic compounds** : Sublimation, Crystallisation, Distillation, Differential extraction and Chromatography.
2. **Chromatography** : It is an important technique used to separate mixtures into their components, purify compounds and also to test the purity of compounds.
3. **Lassaigne test** : It is a test used to detect the presence of N, S, P and X in organic compound.

1.	L.S + FeSO <sub>4</sub> + con H <sub>2</sub> SO <sub>4</sub>	Prussian blue colour	N - present
2.	L.S + CH <sub>3</sub> COOH + (CH <sub>3</sub> COO) <sub>2</sub> Pb	Black ppt	S - present
3.	L.S + Sodium nitrogen prusside	violet colour	S - present
4.	L.S + HNO <sub>3</sub> + AgNO <sub>3</sub>	White ppt Soluble in NH <sub>4</sub> OH	Cl - present
		yellowish ppt. Sparingly soluble in NH <sub>4</sub> OH	Br - present
		yellow ppt Insoluble in NH <sub>4</sub> OH	I - present
5.	L. S + con HNO <sub>3</sub> + ammonium molybdate	yellow ppt	P - present

L.S = Lassaigne solution. It is also called sodium fusion extract.

$$4. \text{ Percentage of } C \text{ (Leibig's method)} = \frac{12 \times \text{mass of } CO_2 \times 100}{44 \times \text{mass of O.C.}}$$

$$5. \text{ Percentage of } H \text{ (Leibig's method)} = \frac{2 \times \text{mass of } H_2O \times 100}{18 \times \text{mass of O.C.}}$$

$$6. \text{ Percentage of } N \text{ (Duma's method)} = \frac{28 \times \text{vol of } N_2 \text{ at S.T.P.}}{224 \text{ mass of O.C.}}$$

$$7. \text{ Percentage of } N \text{ (Kjeldahl's method)} = \frac{1.4 \times N \times V}{\text{mass of O.C.}}$$

Where, N = Normality of acid. V = Vol. of acid used.

$$8. \text{ *Percentage of X* (Carius method) = } \frac{\text{At mass of } X \times \text{mass of Ag X} \times 100}{\text{molar mass of Ag X} \times \text{mass of O.C}}$$

$$9. \text{ *Percentage of S* (Carius method) = } \frac{32 \times \text{mass of BaSO}_4 \times 100}{233 \times \text{mass of O.C}}$$

$$10. \text{ *Percentage of P* (Carius method) = } \frac{62 \times \text{mass of Mg}_2\text{P}_2\text{O}_7 \times 100}{222 \times \text{mass of O.C}}$$

$$11. \text{ *Molecular weight of carboxylic acid* = } n \left[ \left( \frac{108 \times \text{mass of silver salt of acid}}{\text{mass of silver}} \right) - 107 \right]$$

Where, n = Basicity of the acid

$$12. \text{ *Molecular weight of organic base* = } \frac{n}{2} \left[ \left( \frac{195 \times \text{mass of platonic chloride}}{\text{mass of platinum}} \right) - 410 \right]$$

Where, n = Acidity of the base

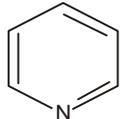
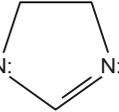
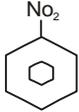
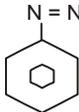
### M.C.Q.

- Which of the following is a technique of purification of organic compound ?  
 (a) Oxidation                      (b) Reduction                      (c) Distillation                      (d) Hydrolysis
- Which of the following is/are method (s) of checking the purity of organic compound ?  
 (a) Chromatography    (b) Spectroscopy                      (c) Both (a) and (b)    (d) Neither (a) nor (b)
- The formula representing a simple ratio of atoms of different elements present in a molecule of the substance is called  
 (a) Empirical formula    (b) Molecular formula    (c) Structural formula    (d) Condensed formula
- What is called the purification technique based on difference in the solubilities of the compound and the impurities in a given solvent ?  
 (a) Sublimation                      (b) Crystallization                      (c) Distillation                      (d) Fractional distillation
- A compound contains C = 90% and H = 10 %. Find empirical formula of the compound.  
 (a) C<sub>3</sub>H<sub>10</sub>                      (b) CH<sub>2</sub>                      (c) C<sub>3</sub>H<sub>2</sub>                      (d) C<sub>3</sub>H<sub>4</sub>
- Empirical formula of a compound is CH<sub>2</sub>O and its vapour density is 30. What is the molecular formula of compound ?  
 (a) C<sub>3</sub>H<sub>6</sub>O<sub>3</sub>                      (b) C<sub>2</sub>H<sub>4</sub>O<sub>2</sub>                      (c) C<sub>2</sub>H<sub>4</sub>O                      (d) CH<sub>2</sub>O
- An organic compound contains C = 36%, H = 6%. What is the empirical formula of compound ?  
 (a) C<sub>3</sub>H<sub>6</sub>                      (b) CH<sub>3</sub>                      (c) CH<sub>2</sub>O                      (d) C<sub>2</sub>H<sub>2</sub>O<sub>2</sub>

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8. What should be the property of solvent used in crystallisation ?  
(a) Impure compound should be sparingly soluble in it at room temperature.  
(b) Impure compound should be appreciably soluble in it at higher temperature.  
(c) Both (a) and (b) (d) Neither (a) nor (b)
9. Which method of purification of organic compound is useful to separate volatile liquids from non volatile impurities ?  
(a) sublimation (b) Distillation  
(c) crystallisation (d) Fractional crystallisation
10. Which method of purification of organic compound is useful to separate liquids having sufficient difference in their boiling point ?  
(a) Sublimation (b) Distillation (c) Filtration (d) Crystallisation
11. An organic compound on analysis gave C = 48gm, H = 8gm and N = 56gm. Volume of 1.0 gm of the compound was found to be 200 ml at STP. What is the molecular formula of the compound ?  
(a)  $C_4H_8N_4$  (b)  $C_2H_4N_2$  (c)  $C_{12}H_{24}N_{12}$  (d)  $C_{16}H_{32}N_{16}$
12. Which of the following method can be employed to separate a mixture of Chloroform (b.p334k) and aniline (b.p 457k) ?  
(a) Distillation (b) Crystallisation (c) Sublimation (d) All of these
13. Insulin contains 3.4% Sulphur. The minimum molecular weight of insulin is  
(a) 350 (b) 470 (c) 560 (d) 940
14. Which element is estimated by Carius method ?  
(a) Nitrogen (b) Carbon (c) Hydrogen (d) Halogen
15. Fractional distillation is used when,  
(a) Difference in melting points of two solids is less.  
(b) Difference in boiling points of two liquids is large.  
(c) Difference in boiling points of two liquids is small.  
(d) Difference in melting points of two liquids is large.
16. Distillation under reduced pressure is used when ...  
(a) liquid has very high boiling point and decompose at or below its boiling point.  
(b) liquid has very low boiling point.  
(c) liquid is insoluble in water.  
(d) liquid is insoluble in organic solvent.
17. In Carius method, 0.099 g compound gave 0.287 gm AgCl. What is the % of Cl ?  
(a) 35.4 (b) 71.7 (c) 46.2 (d) 80.5
18. In which compound, sulphur of organic compound is converted when organic compound is reacted with con.  $HNO_3$  for estimation of Sulphur ?  
(a)  $H_2SO_4$  (b)  $SO_2$  (c)  $H_2S$  (d)  $SO_3$

19. Which of the following techniques is used to separate glycerol from spent-lye in soap industry ?
- (a) Simple distillation (b) Fractional distillation  
(c) Distillation under reduced pressure (d) None of these
20. Which method is used to separate aniline water mixture ?
- (a) steam distillation (b) distillation under reduced pressure  
(c) sublimation (d) None of these
21. Chromatography is used
- (a) to separate mixtures (b) to purify compounds  
(c) to test the purity of compounds (d) All of these
22. In chromatography  $R_f$  means
- (a) Retardation factor (b) Reduced formula  
(c) Reduction factor (d) Retardation frequency
23. Statements regarding Chromatography are given below. Identify True or False statements and give your answer.
- (1) Silica - gel is used as adsorbent in TLC. (2)  $R_f$  = Retardation factor.  
(3) Adsorbent is a mobile phase. (4) Chromaplate is not a part of TLC.
- (a) TTFF (b) TTTT (c) TFFT (d) FFTF
24. In Chromatography, ninhydrin solution is used to detect
- (a) amino acids (b) ethers (c) hydrocarbons (d) glucose
25. In a method of detecting C and H, organic compound is heated with ...
- (a) CuO (b)  $\text{Ca}(\text{OH})_2$  (c)  $\text{CuSO}_4$  (d) All of these
26. In a detection of C and H, the water produced is tested by
- (a) CuO (b)  $\text{Ca}(\text{OH})_2$  (c)  $\text{CuSO}_4$  (d) All of these
27. Lassaigne test is used for detection of
- (a) N (b) S (c) P (d) All of these
28. Lassaigne test is used for detection of
- (a) S (b) P (c) X (d) All of these
29. The basic principle of Lassaigne test is
- (a) to convert elements present in compound from covalent form to ionic form.  
(b) to convert elements present in compound from ionic form to covalent form.  
(c) to convert elements present in compound into their complexes  
(d) None of the above.
30. The prussian blue colour obtained in Lassaigne test is due to
- (a)  $\text{Na}_4[\text{Fe}(\text{CN})_6]$  (b)  $\text{Fe}_4[\text{Fe}(\text{CN})_6]$  (c)  $\text{CuSO}_4$  (d)  $\text{FeSO}_4$

31. In the Lassaigne test of detection of Sulphur, black precipitates are obtained. These are due to the formation of  
 (a)  $\text{Na}_2\text{S}$  (b)  $(\text{CH}_3\text{COO})_2\text{Pb}$  (c)  $\text{PbS}$  (d)  $\text{CH}_3\text{COOH}$
32. What colour is observed when Lassaigne solution of Sulphur containing compound is treated with Sodium Nitroprusside ?  
 (a) Red (b) Violet (c) Blue (d) Brown
33. In a Lassaigne test of Sulphur containing compound the violet colour is due to the  
 (a)  $[\text{Fe}(\text{CN})_5\text{NO}]^{2-}$  (b)  $[\text{Fe}(\text{CN})_5\text{NOS}]^{4-}$  (c)  $\text{Na}_4[\text{Fe}(\text{CN})_6]$  (d)  $\text{K}_4[\text{Fe}(\text{CN})_6]$
34. During the estimation of C and H, the  $\text{CO}_2$  produced is absorbed in  
 (a)  $\text{KOH}$  (b)  $\text{An. CaCl}_2$  (c)  $\text{CuO}$  (d)  $\text{O}_2$
35. During the estimation of C and H, the  $\text{H}_2\text{O}$  produced is absorbed in  
 (a)  $\text{KOH}$  (b)  $\text{An. CaCl}_2$  (c)  $\text{CuO}$  (d)  $\text{O}_2$
36. In Kjeldahl's method for the estimation of nitrogen, the formula used is  
 (a)  $\%N = \frac{1.4NV}{W}$  (b)  $\%N = \frac{NV}{1.8}$  (c)  $\%N = \frac{1.4VW}{N}$  (d)  $\%N = \frac{1.4NW}{V}$
37. An organic compound on analysis gives the following result : C = 54.5%, O = 36.4%, H = 9.1%. The empirical formula of the compound is  
 (a)  $\text{CH}_3\text{O}$  (b)  $\text{C}_2\text{H}_4\text{O}$  (c)  $\text{C}_3\text{H}_4\text{O}$  (d)  $\text{C}_4\text{H}_8\text{O}$
38. An organic compound gave the following results on analysis. C = 53.3%, H = 15.6%, N = 31.1%. Find molecular formula of compound. (Molecular Weight = 45)  
 (a)  $\text{C}_2\text{H}_5\text{N}_2$  (b)  $\text{C}_2\text{H}_7\text{N}$  (c)  $\text{C}_2\text{H}_5\text{N}$  (d)  $\text{C}_2\text{H}_6\text{N}$
39. An organic compound contains C, H and O in the proportion of 6 : 1 : 8 by weight, respectively. Find its molecular formula (V.D = 30)  
 (a)  $\text{CH}_2\text{O}$  (b)  $\text{C}_3\text{HO}$  (c)  $\text{C}_2\text{H}_4\text{O}_2$  (d)  $\text{CH}_4\text{O}$
40. 64 gm of organic compound contains 24 gm C, 8 gm H and the rest oxygen. What is the empirical formula of the compound ?  
 (a)  $\text{CH}_2\text{O}$  (b)  $\text{CH}_4\text{O}$  (c)  $\text{C}_2\text{H}_4\text{O}$  (d)  $\text{C}_2\text{H}_8\text{O}_2$
41. 0.2595 gm of organic substance gave 0.35 gm  $\text{BaSO}_4$ . Find % S in the substance.  
 (a) 18.5 (b) 182 (c) 17.5 (d) 175
42. Find the incorrect formula  
 (a)  $\%C = \frac{12 \times m \times 100}{44 \times V}$  (b)  $\%C = \frac{12 \times m_2 \times 100}{44 \times m_1}$   
 (c)  $\%H = \frac{2 \times m_2 \times 100}{18 \times m_1}$  (d)  $\%N = \frac{28 \times V \times 100}{22400 \times m}$

43. On complete combustion, 0.246g organic substance gave 0.198 g CO<sub>2</sub> and 0.1014 g H<sub>2</sub>O. Find % of C and H in compound ?  
 (a) 4.58, 21.95 (b) 21.95, 4.58 (c) 19, 21 (d) 21, 19
44. Which of the following is the method of estimating nitrogen in organic Compound ?  
 (a) Dumas method (b) Kjeldahl's method (c) Both (A) & (B) (d) Neither(a) nor (b).
45. In the Dumas method nitrogen containing organic compound is first heated with  
 (a) O<sub>2</sub> (b) CuO (c) CO<sub>2</sub> (d) CuO and CO<sub>2</sub>
46. In Dumas method the gaseous mixture produced is passed through KOH<sub>(aq)</sub>. As a result of this \_\_\_\_\_ is absorbed in KOH<sub>(aq)</sub>.  
 (a) CO<sub>2</sub> (b) O<sub>2</sub> (c) NO<sub>x</sub> (d) All of these
47. A device called nitrometer is used in  
 (a) Dumas method (b) Kjeldahl's method (c) Carius method (d) Leibig's method
48. In Dumas method the formula of % N is  
 (a)  $\frac{28 \times m \times 100}{22400 \times v}$  (b)  $\frac{28 \times V \times 100}{22400 \times m}$  (c)  $\frac{14 \times V \times 100}{22400 \times m}$  (d)  $\frac{14 \times m \times 100}{22400 \times v}$
49. In Dumas method for estimation of nitrogen, 0.3g of an organic compound gave 50 ml of nitrogen collected at 300K temperature and 715 mm pressure. Calculate the % of N in the compound. (Aqueous tension at 300K = 15mm)  
 (a) 20.8 % (b) 17.46 % (c) 34.5% (d) 25 %
50. In Kjeldahl's method, when nitrogenous organic compound is heated with con. sulphuric acid, the nitrogen is converted into  
 (a) (NH<sub>4</sub>)<sub>2</sub> SO<sub>4</sub> (b) NH<sub>4</sub>OH (c) NH<sub>3</sub> (d) None
51. Kjeldahl's method cannot be applied to  
 (a)  (b)  (c)  (d) None of these
52. Kjeldahl's method, can be applied to  
 (a) CH<sub>3</sub>CH<sub>2</sub>NH<sub>2</sub> (b) CH<sub>3</sub>CH<sub>2</sub>NO<sub>2</sub> (c)  (d) None of these
53. During estimation of nitrogen present in organic compound by Kjeldahl's method, the ammonia evolved from 0.5 gm of the compound neutralized 10 ml of 1 M H<sub>2</sub>SO<sub>4</sub>. Calculate the %N in the compound.  
 (a) 65 (b) 28 (c) 56 (d) 30
54. In the estimation of halogen by Carius method, the halogen of the organic compound is converted into  
 (a) Silver nitrate (b) Silver halide (c) Silver Sulphate (d) Silver oxide

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55. In Carius method of estimation of halogen, 0.15 gm of an organic compound gave 0.12gm of AgBr. Calculate % of bromine in the compound.  
(a) 34.04%                      (b) 17.02%                      (c) 53.19%                      (d) 18.8%
56. In estimation of Sulphur in organic compound, the amount of sulphur is determined in terms of  
(a)  $H_2SO_4$                       (b)  $BaSO_4$                       (c)  $SO_2$                       (d)  $SO_3$
57. In estimation sulphur, 0.157gm of an organic compound gave 0.4813 gm of barium Sulphate. Calculate % S in compound.  
(a) 24.10 %                      (b) 14.20%                      (c) 42.10%                      (d) 38%
58. In estimation of phosphorous present in organic compound, the amount of phosphorous is determined in terms of  
(a) Phosphoric acid                      (b) Ammonium molybdate  
(c) Ammonium phospho molybdate                      (d) Magnesium pyrophosphate
59. Suggest a method to purify camphor containing impurity of common salt.  
(a) Crystallisation                      (b) Fractional crystallisation  
(c) Sublimation                      (d) Distillation
60. Suggest a method to separate a mixture of ortho and para nitro aniline.  
(a) Distillation                      (b) Chromatography  
(c) Differential extraction                      (d) All of these
61. Suggest a method to separate a mixture of methanol and acetone.  
(a) Simple distillation                      (b) Fractional distillation  
(c) Fractional Crystallisation                      (d) Differential extraction
62. Suggest a method to purify naphthalene.  
(a) Sublimation                      (b) Crystallisation  
(c) Differential extraction                      (d) Distillation
63. Suggest a method to separate a mixture containing acetone and ethanol  
(a) Sublimation                      (b) Crystallisation  
(c) Differential extraction                      (d) Distillation
64. Which of the following compound will not give positive Lassaigne test for nitrogen ?  
(a) Hydrazine                      (b) urea                      (c) ethylamine                      (d) nitroethane.
65. Name a suitable technique of separation of a mixture of calcium sulphate and camphor.  
(a) Crystallisation                      (b) Sublimation                      (c) Distillation                      (d) None
66. The fragrance of flower is due to the presence of some steam volatile organic compounds called essential oils. These are generally insoluble in water at room temperature but are miscible with water vapour. A suitable method for the extraction of these oils from the flower is,  
(a) Distillation                      (c) Crystallisation  
(c) Distillation under reduced pressure                      (d) Steam distillation

67. Kjeldahl's method can not be used for estimation of nitrogen in ...  
 (a) DNA (b) RNA (c) a and B (d) Neither a and b
68. Name the method used to separate a mixture of iodine and sodium chloride ?  
 (a) Crystallisation (b) Fractional Crystallisation  
 (c) Sublimation (d) Steam distillation
69. What conclusion can be drawn if red colour is obtained during Lassaigne test ?  
 (a) Pressure of S (b) Pressure of N  
 (c) both a and b (d) Neither a and b
70. What is the formula of iron (III) hexacyanoferrate (II) ?  
 (a)  $\text{Fe}_4[\text{Fe}(\text{CN})_6]_3$  (b)  $\text{Fe}_3[\text{Fe}(\text{CN})_6]_4$  (c)  $\text{Fe}_4[\text{Fe}(\text{CN})_6]$  (d)  $\text{Fe}_3[\text{Fe}(\text{CN})_6]$
71. Two volatile liquids A and B differ in their boiling points by 15K. Which process can be used for their separation ?  
 (a) Fractional distillation (b) Steam distillation  
 (c) Distillation under reduced pressure (d) Simple distillation
72. An organic compound from its aqueous solution can be separated by  
 (a) Distillation (b) Solvent extraction  
 (c) Fractional distillation (d) Steam distillation
73. The ammonia evolved from the treatment of 0.3gm of organic compound for estimation of nitrogen was passed in 100ml of 0.1M  $\text{H}_2\text{SO}_4$ . The excess acid required 20ml 0.5 M NaOH for complete neutralisation. The organic compound is...  
 (a) acetamide (b) benzamide (c) urea (d) thiourea
74. Which of the following compound gives blood red colouration when its Lassaigne extract is treated with alkali and ferric chloride ?  
 (a) Thiourea (b) Diphenyl Sulphide  
 (c) Phenyl hydrazine (d) Benzamide
75. Which of the following compounds will give Lassaigne test for nitrogen ?  
 (a)  $\text{NH}_2\text{NH}_2$  (b)  $\text{NaNO}_3$  (c)  $\text{NH}_4\text{Cl}$  (d) NaCN
76. Which of the following will give bloodred colour while doing Lassaigne test for nitrogen  
 (a) urea (b) Benzene sulphonic acid  
 (c) Hydrazine (d) p-aminobenzene sulphonic acid
77. Name the method used to determine the molecular weight of organic compound.  
 (a) Silver salt method (b) Chloroplatinate salt method  
 (c) Mass spectrometry (d) All of these
78. While detecting phosphorous in organic compound yellow colour is obtained due to the formation of  
 (a)  $(\text{NH}_4)_2\text{MoO}_4$  (b)  $(\text{NH}_4)_3\text{PO}_4$  (c)  $(\text{NH}_4)_3\text{PO}_4 \cdot 12\text{MoO}_3$  (d)  $\text{MgNH}_4\text{PO}_4$

79. 0.306 gm of the silver salt of a monobasic organic acid yields 0.216gm of Silver on ignition. Find molecular weight of the acid.  
 (a) 46 (b) 23 (c) 122 (d) 20
80. 1.026gm of the silver salt of a tribasic organic acid gave an ignition 0.648gm silver. Find molecular weight of the acid.  
 (a) 148 (b) 192 (c) 208 (d) 234
81. 0.3557gm of the Platinum salt of monoacidic organic base yields 0.117gm of platinum. Find molecular weight of base  
 (a) 95 (b) 91.45 (c) 98 (d) 410
82. Chromatography is used for the purification of  
 (a) solids (b) liquids (c) gases (d) All of these
83. Which of the following reagents is used for detection of unsaturation in alkenes.  
 (a) NaOH + CaO (b) cold dil, alkaline  $\text{KMnO}_4$   
 (c)  $\text{Cl}_2/\text{hv}$  (d) KOH / EtOH
84. The presence of halogen in an organic compound is detected by  
 (a) Iodo form test (b) Silver mirror test  
 (c) Beilstein's test (d) Millon's test
85. Which of the following pairs has the same percentage of carbon ?  
 (a)  $\text{CH}_3\text{COOH}$  &  $\text{C}_2\text{H}_5\text{OH}$  (b)  $\text{C}_6\text{H}_{12}\text{O}_6$  &  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$   
 (c)  $\text{HCOOCH}_3$  &  $\text{C}_2\text{H}_5\text{O}$  (d)  $\text{CH}_3\text{COOH}$  &  $\text{C}_6\text{H}_{12}\text{O}_6$
86. An organic compound having molecular mass 60 is found to contain C=20%, H=6.67% and N=46.67% while rest is oxygen. On heating it gives  $\text{NH}_3$  along with a solid residue. The solid residue give violet colour with alkaline  $\text{CuSO}_4$ , the compound is  
 (a)  $\text{CH}_3\text{NCO}$  (b)  $\text{CH}_3\text{CONH}_2$  (c)  $(\text{NH}_2)_2\text{CO}$  (d)  $\text{CH}_3\text{CH}_2\text{CONH}_2$
87. How will you separate a solution of benzene and chloroform ?  
 (a) Sublimation (b) Filtration (c) Distillation (d) Crystallisation
88. When 32.25 gm ethyl chloride dehydrohalogenated it gives 50% alkene. What is the mass of product ?  
 (a) 7 gm (b) 14 gm (c) 21 gm (d) 28gm
89. How much sulphur is present in organic compound if on analysis 0.53gm of this compound gives 1.158gm of  $\text{BaSO}_4$  ?  
 (a) 10% (b) 20% (c) 30% (d) 40%
90. Which method can be used to separate a mixture of KCl and  $\text{KClO}_3$  ?  
 (a) Crystallisation (b) Fractional Crystallisation  
 (c) Sublimation (d) Distillation
91. Middle ton's test is used to detect  
 (a) N (b) P (c) X (d) S

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92. An organic compound weighing 4.337mg yielded 10.35mg of  $\text{CO}_2$  and 3.42mg of  $\text{H}_2\text{O}$  on oxidation with  $\text{CuO}$ . Calculate the empirical formula of the compound.  
(a)  $\text{C}_3\text{H}_5$                       (b)  $\text{C}_3\text{H}_5\text{O}$                       (c)  $\text{C}_4\text{H}_6\text{O}$                       (d)  $\text{C}_{10}\text{H}_{16}\text{O}_3$
93. An organic compound contains 69% C and 4.8% H, the remainder being oxygen. Calculate the masses of  $\text{CO}_2$  and  $\text{H}_2\text{O}$  produced when 0.2gm of this substance is completely combusted.  
(a) 0.506g & 0.864g                      (b) 0.0864g & 0.0506g  
(c) 0.506g & 0.0684g                      (d) 0.506g & 0.0864g
94. 0.3780 gm of chlorine containing organic compound gave 0.5740gm of silver chloride in Carius estimation. Calculate the % of Cl.  
(a) 37.56%                      (b) 40.5%                      (c) 35.7%                      (d) 31.23%
95. In the estimation of sulphur by Carius method, 0.468gm of organic compound gave 0.668gm  $\text{BaSO}_4$ . Calculate % of S.  
(a) 19.6                      (b) 21.7                      (c) 17.9                      (d) 16.9

***In each of the following questions, a statement (A) is given followed by Reason (R). Give your answers as***

- (A) If both A and R are true, and reason is the true explanation of the assertion.  
(B) If both A and R are true, but R is not the true explanation of the A.  
(C) If A is true but R is false.  
(D) If both A and R are false.
96. A: Vapour density and density both are same  
R: Molecular weight is twice the density
97. A: Lassaigne's extract is boiled with dil.  $\text{HNO}_3$  before testing for halogens by  $\text{AgNO}_3$ .  
R:  $\text{CN}^-$  and  $\text{S}^{2-}$  present in Lassaigne extract interfere with the test of halide ions by silver nitrate.
98. A: Lassaigne's test is shown by diazonium salts.  
R: Diazonium salts are volatile.
99. A: Essential oils are purified by steam distillation.  
R: Essential oils are volatile and are insoluble in water.
100. A: Sulphur present in organic compound can be estimated quantitatively by Carius method.  
R: Sulphur is separated easily from other atoms in the molecules and gets precipitated as light yellow solid.

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**Answer Key**

1	C	26	C	51	D	76	D
2	C	27	D	52	A	77	D
3	A	28	D	53	C	78	C
4	B	29	A	54	B	79	A
5	D	30	B	55	A	80	B
6	B	31	C	56	B	81	B
7	C	32	B	57	C	82	D
8	C	33	B	58	D	83	B
9	B	34	A	59	C	84	C
10	B	35	B	60	B	85	D
11	A	36	A	61	B	86	C
12	A	37	B	62	A	87	C
13	D	38	B	63	D	88	A
14	D	39	C	64	A	89	C
15	C	40	B	65	B	90	B
16	A	41	A	66	D	91	D
17	B	42	A	67	C	92	D
18	A	43	B	68	C	93	D
19	C	44	C	69	C	94	A
20	A	45	D	70	A	95	A
21	D	46	A	71	A	96	D
22	A	47	A	72	B	97	A
23	A	48	B	73	C	98	D
24	A	49	B	74	A	99	A
25	A	50	A	75	D	100	C

## HINTS

11. 200ml of compound = 1 gm  $\therefore$  22400 ml of compound = 112 gm

17.  $\%Cl = \frac{35.5}{143.5} \times \frac{0.287}{0.099} \times 100$

39. Elements	No. of moles	Simple ratio
C	6/12 = 0.5	1
H	1/1 = 1	2
O	8/16 = 0.5	1

$\therefore$  E.F = CH<sub>2</sub>O

M.W = 2VD = 60

$\therefore$  n = 2  $\therefore$  M.F = C<sub>2</sub>H<sub>4</sub>O<sub>2</sub>

41.  $\%S = \frac{32}{233} \times \frac{0.35}{0.2595} \times 100 = 18.52$

49. P = 715 - 15 = 700 mm

$Vol\ of\ N_2\ at\ STP = \frac{273 \times 700 \times 50}{300 \times 760} = 41.9\ ml$   $\therefore \%N = \frac{28 \times 41.9 \times 100}{22400 \times 0.3} = 17.46\%$

53. 1M of 10 ml H<sub>2</sub>SO<sub>4</sub> = 1 M of 20ml NH<sub>3</sub>  
1000ml of 1M NH<sub>3</sub> = 14 g nitrogen

20 ml of 1M NH<sub>3</sub> =  $\frac{14 \times 20}{1000}$  g nitrogen

$\%N = \frac{14 \times 20 \times 100}{1000 \times 0.5} = 56.0\%$

73. let the vol of acid left unused = V ml of 0.1 MH<sub>2</sub>SO<sub>4</sub>

$\therefore V \times 0.1 \times 2 = 20 \times 0.5 \times 1$

$\therefore V = 50\ ml = 50\ ml\ of\ .01\ m\ H_2SO_4$

$\therefore Vol\ of\ acid\ used = 100 - 50$

$\%N = \frac{1.4 \times 2 \times 50 \times 0.1}{0.3} = 46.6$

$\%N\ in\ urea = \frac{28}{60} \times 100 = 46.6$

79.  $MW = \frac{108 \times 0.306}{0.216} - 107 = 46$

81.  $MW = \frac{1}{2} \left[ \left( \frac{195 \times 0.3557}{0.117} \right) - 410 \right]$